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Chapter 6 Solutions Engineering and Chemical Thermodynamics Wyatt Tenhaeff Milo Koretsky Department of Chemical Engineering Oregon State University 2 6.1 (a) The Clausius-Clapeyron equation: $dP_i^{sat} / P_i^{sat} = h_i^{vap} dT / RT^2$ or $\ln P_i^{sat} / 101 \text{ kPa} = -h_i^{vap} / R (1/T - 1/373 \text{ [K]})$ so $P_i^{sat} = 101 \text{ kPa} \exp(-h_i^{vap} / R (1/T - 1/373 \text{ [K]}))$

Chapter 6 Solutions - Chapter 6 Solutions Engineering and ...

Chapter 6 2 If the fluid is in thermodynamic equilibrium any thermodynamic variable for a pure substance, like pure water, can be written in terms of any two other thermodynamic variables, i.e. $p=p(\rho,T)$ (6.1.1) where the functional relationship in depends on the substance.

Chapter 6 Thermodynamics and the Equations of Motion

Chapter 6: Solution Thermodynamics and Principles of Phase Equilibria In all the preceding chapters we have focused primarily on thermodynamic systems comprising pure substances. However, in all of nature, mixtures are ubiquitous.

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