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mass of copper in 250.0 cm^3 $\ll 0.0080 \text{ mol dm}^{-3} \times 0.2500 \text{ dm}^3$
 $63.55 \text{ g mol}^{-1} \gg 0.127 \text{ g}$ OR mass of
brass in 1 dm^3 $\ll 4 \times 0.200 \text{ g} \gg 0.800 \text{ g}$
AND $[\text{Cu}^{2+}] = \ll 0.0080 \text{ mol dm}^{-3} \gg$

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$\text{g mol}^{-1} \gg 0.5084 \text{ g dm}^{-3} \ll \% \text{ copper in this sample of brass } 0.127 \text{ } 0.200$

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